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charge. electron. The smallest known particles that make up protons and neutrons. quarks. The positively charged center of that atom where most of the mass of the atom is located. nucleus. The negatively charged area around the nucleus where the volume of the atom is

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Answers

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STUDY GUIDE ANSWERS ATOMIC

STRUCTURE - PERIODIC TABLE TEST 1.

How many periods make up the

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periodic table? 7 2. How many groups or families make up the periodic table? 9 families (The Lanthanide Series and Actinide Series are part of the transition metals family.) There are 18 groups.

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Name _____
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est Date _____. Atomic Structures Study Guide ANSWERS. 1. The Periodic Table is mostly made of metals. 2. The noble gases are the

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Answers
farthest group to the right on the Periodic Table. 3. Atoms having equal numbers of positive and negative charges make them neutral.

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ANSWERS

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Calculate the average atomic mass for Ne given the following isotopes & their abundance: Neon-20 = 90.92%; Neon-21 = 0.257%; Neon-22 = 8.82%. Study Guide Answer Key. Answer the following questions regarding the discovery of the atom & its parts. A. Identify 3 conclusions of the gold foil

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atomic number. 24. d 25. a. 26. d. 27.
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Study Guide – Chapter 4 – The Structure of the Atom
Atomic Structure Study Guide Answers Atomic Structures Study Guide ANSWERS 1. The Periodic Table is mostly made of metals. 2. The

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noble gases are the farthest group to the right on the Periodic Table. 3. Atoms having equal numbers of positive and negative charges make them neutral. 4.

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Answers

1. Describe the evolution of atomic theory leading to the current model of the atom based on the works of Democritus, Dalton, Thomson, Rutherford, Bohr, and Schrodinger.
2. Explain the law of conservation of mass, the law of definite proportions,

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Answers and the law of multiple proportions.

3. Describe the properties of protons, neutrons, and electrons.

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reference see “ Chemistry: An Atoms-focused Approach ” by Gilbert, Kirss, and Foster Chapter 3: Atomic Structure, Explaining the Properties of Elements Trends to know (and be able to explain the trend...think about Z_{eff}): o Effective Nuclear Charge (Z_{eff}): the attraction toward

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Answers experienced by

Chapter 3: Atomic Structure,
Explaining the Properties of ...

Atoms are the smallest unit of matter that cannot be divided using any chemical method. They do consist of

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Smaller parts, but can only be broken by nuclear reactions. The three parts of an atom are protons, neutrons, and electrons. Protons carry a positive electrical charge. Neutrons are electrically neutral.

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Answers Atomic Structure and Atomic Theory - Study Guide

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ATOMIC STRUCTURE: SUBATOMIC

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ANSWERS
PARTICLE CHARGE LOCATION WITHIN THE ATOM ATOMIC MASS (AMU)

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Answers

$x = \text{fraction } ^{37}\text{Cl}$ $1 - x = \text{fraction } ^{35}\text{Cl}$. Instead of using the integers 37 and 35 as atomic masses, take the more precise atomic masses of the isotopes from Table 1: The calculation reveals that natural chlorine is 24% ^{37}Cl and 76% ^{35}Cl .

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