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the shape of your desk; you know that your backpack cannot hold seven textbooks.

You encounter liquids throughout the day as you

Chapter 13: States of Matter

Chapter 13 States of Matter 137

SECTION 13.1 THE NATURE OF

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GASES (pages 385 – 389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. **Kinetic Theory and a Model for Gases (pages 385 – 386)**

1.

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Name _____ Date _____ Class _____ STATES OF
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There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are solid, liquid, and gas. These

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three states are quite different. The main difference is in their particles.

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by Anna

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of the particles of a substance.

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PLAY. Kinetic Molecular Theory.

Explains the properties of gases in terms of the energy, size, and motion of their particles. Elastic Collision. Describes a

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collision in which kinetic energy may be transferred between colliding particles but the total kinetic energy of the two particles remains the same.

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all matter consists of tiny particles that are constantly in motion What are the three assumptions of the kinetic theory as it applies to gases? -The particles in a gas are considered to be small, hard spheres with

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an insignificant volume. - The motion of the particles in a gas are rapid, constant, and random.

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Chapter 13 States of Matter 2 Kinetic Theory as Applied to Gases Fundamental assumptions about gases. The particles in a gas are considered to be small, hard spheres with an insignificant volume. Between particles in a gas there is empty space. No attractive or repulsive forces exist between the particles. 3

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Put this matter into our own hands by writing
and sending #BlackLoveLetters through
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