

## Chemical Quantities Chapter 10 Answers

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8.0 x 10<sup>1</sup> moles. How many atoms are in 0.075 mol of titanium? 4.5 x 10<sup>22</sup>. How many molecules are in 2.10 mol CO<sub>2</sub>? 1.26 x 10<sup>24</sup> molecules. Butanol is composed of carbon, hydrogen, and oxygen. If 1.0 mol of butanol contains 6.0 x 10<sup>24</sup> atoms of hydrogen, what is the subscript for the hydrogen atom in C<sub>4</sub>H<sub>2</sub>O? 10.

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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

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10. Look at Figure 10.16 and Table 10.3. Name three compounds that have an empirical formula of CH.  
11. Fill in the labels on the diagram below. MICROSCOPIC INTERPRETATION MACROSCOPIC INTERPRETATION SO<sub>3</sub> molecule 1 mol SO<sub>3</sub> SO<sub>3</sub> composed of composed of S atom and sulfur atoms ( 1023) oxygen atoms 3 atoms and CHAPTER 10, Chemical Quantities(continued)

~~SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)~~

Answers Chapter 10 (Chemical Quantities) Test Study Guide The mole is Page 11/26. File Type PDF Chemical Quantities Answers Key Chapter Test the SI unit used to measure the number of representative particles in a substance. A representative particle can Page 9/27.

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