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The stoichiometry of  
a reaction describes

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the relative amounts of reactants and products in a balanced chemical equation. A stoichiometric quantity of a reactant is the amount necessary to react completely with the other reactant (s). If a reactant remains unconsumed after complete reaction has

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occurred, it is in  
excess.

~~3: Stoichiometry of  
Formulas and  
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...

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of quantitative  
relationships between  
the amounts of

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reactants used and the amounts of products formed in a chemical reaction.

Conduct two separate mass-to-mass (or mole-to-mole) stoichiometric calculations; comparing the mass of each reactant to a chosen product ...

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~~Stoichiometry ...~~

Stoichiometry

involves the concept  
of the mole.

Particulate matter  
and chemical change  
Writing formulae and  
chemical equations is  
a fundamental skill of  
a chemist (it is also a

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universal language).  
You are probably familiar with balancing chemical equations and many of the key terms in this part of the course.

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In ideal Stoichiometry

– All reactants are  
converted into

products. Common

Methods for solving

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all Stoichiometry

Problems: Mass-Mass

Problems: 1. Start

with a known mass of

reactant or product

and find an unknown

mass of another

reactant or product.

2. All other

stoichiometry

problems are

derivations

(shortened versions)

of this larger solution:

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Stoichiometry is established from the law of conservation of mass wherein it correlates the reactants and the products quantitatively.

Chemical reactions have a reactant side and product side.

The... Why is

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"Chemistry is so crucial to an understanding of medicine and biology, environmental science, and many areas of engineering



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