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Aqueous Solutions and Precipitation Reactions

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Note: Not all of the reactions will occur.

For those that do not, write no reaction. 1.

A piece of copper is dropped into a container

of water. No reaction 2. Liquid bromine is

added to a container of sodium iodide

crystals. $\text{Br (l)} + 2\text{NaI(s)} \rightarrow 2\text{NaBr(s)} + \text{I}_2\text{(s)}$

3. An aluminum strip is immersed in a

solution of silver nitrate.

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tin(II) carbonate sodium hydrogen carbonate
NaHCO 1. 2. 3. 4. 5. 6. 8. 9. 10. 11.
vanadium(V) oxide H₂O dihydrogen monoxide
(NH)₄CO ammonium oxalate 4 224 polonium(VI)
thiocyanate Po(SCN)₆ tetraphosphorus

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decaoxide P O 4 10 12. 13. 15. 16. 17. 18.
19. manganese(VII) oxide copper(II)
dihydrogen phosphate

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Since many of the transuranium elements were prepared at the Lawrence Radiation Center, University of California at Berkeley, California, U.S.A., it is logical to find the following names: Americium—named after the Americas (after all, one must be patriotic!).

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